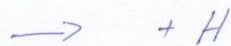
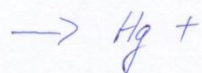


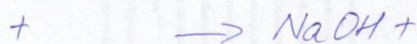
1. Kolik g kyslíku vznikne rozkladem 1g vody elektrickým proudem?



2. Kolik g kyslíku vznikne kalcinací 5g oxidu rtuťnatého (HgO)



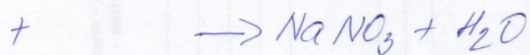
3. Kolik g sodíku vznikne, použijeme-li k rozkladu vody 5g sodíku.



4. Kolik g síry shori v 1g kyslíku?

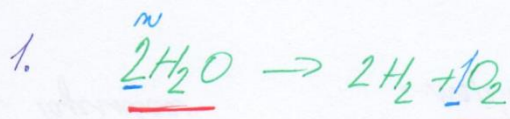


5. Kolik g vody vznikne při neutralizaci 1g kyseliny dusičné (HNO_3) hydroxidem sodným ($NaOH$)?



DOPLNĚNO O ROVNICI

VÝSLEDKY - VÝPOČTY HMOTNOSTI Z CHEM. ROVNIC



$m = n \cdot M$

$$m(\text{H}_2\text{O}) = 2 \cdot (10 \cdot 2 + 16) = 36,04 \text{ g}$$

$$m(\text{O}_2) = 1 \cdot 16 \cdot 2 = 32 \text{ g}$$

úměrní k též

$$\begin{array}{rcl} \uparrow 36,04 \text{ g H}_2\text{O} & \dots & 32 \text{ g O}_2 \uparrow \\ 1 \text{ g} & & x \end{array}$$

$$x = \underline{\underline{0,89 \text{ g}}}$$

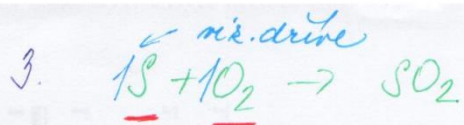


$$m(\text{HgO}) = 2 \cdot (200,59 + 16) = 433,18 \text{ g}$$

$$m(\text{O}_2) = 16 \cdot 2 = 32 \text{ g}$$

$$\begin{array}{rcl} \uparrow 2 \cdot 433,18 \text{ g HgO} & \dots & 32 \text{ g O}_2 \uparrow \\ \uparrow 2 \cdot 5 \text{ g HgO} & & x \text{ g O}_2 \end{array}$$

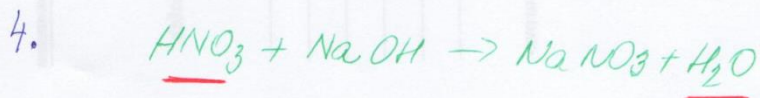
$$x = \underline{\underline{0,37 \text{ g}}}$$



$$m(S) = 32,06g$$

$$m(O_2) = 16 \cdot 2 = 32g$$

$$\begin{array}{r} \uparrow \times 32,06g S \dots 32g O_2 \uparrow \\ \times 1g \dots \times \\ \hline x = \underline{\underline{1g}} \end{array}$$



$$m(HNO_3) = 1,01 + 14,01 + 3 \cdot 16 = 63,02g$$

$$m(H_2O) = 1,01 \cdot 2 + 16 = 18,02g$$

$$\begin{array}{r} \uparrow \times 63,02g HNO_3 \dots 18,02g H_2O \uparrow \\ \times 1g HNO_3 \dots \times \\ \hline x = \underline{\underline{0,29g}} \end{array}$$